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D. Keffler, Dept. of Chemical Engineering, University of Tennessee, Knoxville, February 19, 2001 1 I. CALCULATING THE EQUILIBRIUM COEFFICIENT FROM PHYSICAL PROPERTIES A. Given information: 1. Thermodynamic state of system temperature T pressure p 2. Reaction Stoichiometry v where v_{ij} is the stoichiometric coefficient of species j in reaction i.

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Second, we have mechanical equilibrium, equating the pressure in the two phases, $PV = \mu$. (2) Third, we have chemical equilibrium, equating the chemical species of each component in the two phases, $V L \mu_a = \mu_a$ for $a = 1$ to N_c , (3) where N_c is the number of components in the mixture. II. Equilibrium Expression in terms of Gibbs free energies

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